

## UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS

International General Certificate of Secondary Education

### MARK SCHEME for the June 2004 question papers

#### 0620 CHEMISTRY

0620/01	Paper 1 (Multiple Choice), maximum mark 40
0620/02	Paper 2 (Core), maximum mark 80
0620/03	Paper 3 (Extended), maximum mark 80
0620/05	Paper 5 (Practical), maximum mark 40
0620/06	Paper 6 (Alternative to Practical), maximum mark 60

These mark schemes are published as an aid to teachers and students, to indicate the requirements of the examination. They show the basis on which Examiners were initially instructed to award marks. They do not indicate the details of the discussions that took place at an Examiners' meeting before marking began. Any substantial changes to the mark scheme that arose from these discussions will be recorded in the published *Report on the Examination*.

All Examiners are instructed that alternative correct answers and unexpected approaches in candidates' scripts must be given marks that fairly reflect the relevant knowledge and skills demonstrated.

Mark schemes must be read in conjunction with the question papers and the *Report on the Examination*.

- CIE will not enter into discussion or correspondence in connection with these mark schemes.

CIE is publishing the mark schemes for the June 2004 question papers for most IGCSE and GCE Advanced Level syllabuses.



**Grade thresholds** taken for Syllabus 0620 (Chemistry) in the June 2004 examination.

	maximum mark available	minimum mark required for grade:			
		A	C	E	F
Component 1	40	-	26	20	17
Component 2	80	-	52	36	27
Component 3	80	53	31	-	-
Component 5	40	31	24	18	14
Component 6	60	42	32	21	15

The threshold (minimum mark) for B is set halfway between those for Grades A and C.  
The threshold (minimum mark) for D is set halfway between those for Grades C and E.  
The threshold (minimum mark) for G is set as many marks below the F threshold as the E threshold is above it.

Grade A\* does not exist at the level of an individual component.

**June 2004**

**INTERNATIONAL GCSE**

**MARK SCHEME**

**MAXIMUM MARK: 40**

**SYLLABUS/COMPONENT: 0620/01**

**CHEMISTRY  
(Multiple Choice)**



<b>Page 1</b>	<b>Mark Scheme</b>	<b>Syllabus</b>	<b>Paper</b>
	<b>Chemistry – June 2004</b>	<b>0620</b>	<b>01</b>

<i>Question Number</i>	<i>Key</i>	<i>Question Number</i>	<i>Key</i>
1	<b>A</b>	21	<b>C</b>
2	<b>D</b>	22	<b>C</b>
3	<b>B</b>	23	<b>B</b>
4	<b>B</b>	24	<b>D</b>
5	<b>C</b>	25	<b>D</b>
6	<b>C</b>	26	<b>A</b>
7	<b>A</b>	27	<b>B</b>
8	<b>D</b>	28	<b>B</b>
9	<b>A</b>	29	<b>C</b>
10	<b>D</b>	30	<b>C</b>
11	<b>A</b>	31	<b>D</b>
12	<b>B</b>	32	<b>A</b>
13	<b>B</b>	33	<b>A</b>
14	<b>D</b>	34	<b>B</b>
15	<b>C</b>	35	<b>A</b>
16	<b>D</b>	36	<b>D</b>
17	<b>B</b>	37	<b>A</b>
18	<b>C</b>	38	<b>D</b>
19	<b>A</b>	39	<b>B</b>
20	<b>A</b>	40	<b>A</b>

**June 2004**

**INTERNATIONAL GCSE**

**MARK SCHEME**

**MAXIMUM MARK: 80**

**SYLLABUS/COMPONENT: 0620/02**

**CHEMISTRY**



Page 1	Mark Scheme	Syllabus	Paper
	Chemistry - June 2004	0620	02

- 1 (a) B, C, F (all needed); [1]  
 Only contain one type of atom [1]  
 NOT: contain one kind of molecule  
 NOT: cannot be split using chemical means
- (b) C [1]
- (c) (i) B [1]  
 (ii) any gas with diatomic molecules e.g. chlorine, hydrogen, hydrogen chloride [1]
- (d) (i) F [1]  
 (ii) pencil 'leads'/in pencils/lubricant/in electrical conductors/for electrodes/  
 in tennis racquets/in golf clubs/hockey sticks etc [1]
- (e) (i) substance containing 2 or more different atoms  
 combined/bonded/joined (both parts needed for mark) [1]  
 ALLOW: elements (chemically) combined  
 (ii) methane [1]
- (f) (i) 8 electrons round chlorine and bonded pair with dot and cross = 2 [2]  
 ALLOW: all dots or all crosses  
 Correct number of electrons but bonded pair not clearly on overlap = 1  
 NOT: molecules other than hydrogen chloride  
 (ii) covalent [1]  
 (iii) blue litmus; [1]  
 (litmus) turns red [1]  
 (iv) pH2 [1]  
 (v) 2 [1]  
 (vi) magnesium chloride [1]  
 NOT: formula
- Total = 17**
- 2 (a) insoluble particles/solids/dirt trapped/caught on stones; [1]  
 NOT: filter reacts with insoluble impurities  
 NOT: impurities unqualified  
 Water passes through/filtered OWTTE [1]
- (b) (i) kill bacteria/germs, disinfect water OWTTE [1]  
 (ii) neutralises acidity/water [1]  
 ALLOW: reacts with acids in water  
 (iii) calcium hydroxide [1]  
 NOT: formula  
 (iv) neutralising acid soils/neutralising acidic (industrial) waste/making  
 bleaching powder/removing acidic gases/in Solvay process/in recovery of  
 ammonia/making limewater/in water softening/for making plaster/for  
 making mortar/controlling soil acidity [1]  
 NOT: neutralising acids unqualified  
 NOT: making cement

Page 2	Mark Scheme	Syllabus	Paper
	Chemistry - June 2004	0620	02

- (c) (i) 100; [1]  
°C (conditional on 100) [1]
- (ii) anhydrous cobalt chloride/anhydrous copper sulphate (or correct colours); [1]  
NOT: cobalt chloride/copper sulphate unqualified  
Turns pink/blue (respectively) [1]
- (iii) any suitable e.g. washing/cleaning/drinking/cooking [1]
- (d) B [1]
- (e) ethanol [1]  
NOT: alcohol
- (f) potassium hydroxide; hydrogen [1]  
NOT: symbols
- Total = 15**
- 3 (a) means of measuring gas volume e.g. gas syringe/measuring cylinder [1]  
(must be graduated);  
flask/test tube/vessel with calcium carbonate + acid leading to syringe etc [1]  
IGNORE: lack of reference to closed system (unless drawing incorrect) [1]  
record volume on gas syringe/measuring cylinder/measure how much  
carbon dioxide given off [1]  
at various time intervals/at a particular time [1]  
OR  
flask/vessel with calcium carbonate and hydrochloric acid in flask (1)  
measure its mass at beginning of experiment (1)  
measure mass of flask and contents during reaction (1)  
at specific time(s) (1)
- (b) (i) faster/greater/speeds up [1]
- (ii) slower/less [1]
- (iii) faster/greater/speeds up [1]
- (c) (i) add aqueous sodium hydroxide; [1]  
white precipitate; [1]  
insoluble in excess [1]  
(incorrect reagent = 0)  
ALLOW: flame test - brick red
- (d) (i) high in the reactivity series/very reactive [1]
- (ii) 2 electrons in outer shell; [1]  
inner shells correct as 2.8.8 [1]
- Total = 13**

Page 3	Mark Scheme	Syllabus	Paper
	Chemistry - June 2004	0620	02

- 4 (a) ethanol - solvent  
ethene - polymer  
bitumen - roads [3]
- (b) ethanol [1]
- (c) (i) C [1]  
(ii) A [1]  
(iii) B [1]  
(iv) D [1]
- (d) (i) (compound) containing only carbon and hydrogen [1]  
NOT: it contains carbon and hydrogen
- (ii) has only single bonds/ has general formula  $C_nH_{2n+2}$  [1]  
NOT: it is saturated

**Total = 10**

- 5 (a) chlorine, argon, potassium, bromine, iodine [1]  
ALLOW: symbols
- (b) chlorine, potassium, argon, bromine, iodine [1]  
ALLOW: symbols
- (c) 2<sup>nd</sup> box down ticked [1]
- (d) chlorine, bromine, iodine (all 3 needed) [1]  
ALLOW: symbols
- (e) (i) potassium/K [1]  
(ii) argon/Ar [1]
- (f) 1<sup>st</sup> and 4<sup>th</sup> boxes ticked (1 mark each) [2]
- (g) (i) high (boiling point) [1]  
(ii) conducts/is high [1]
- (h) potassium loses an/one electron/loses outer shell [1]  
chlorine gains an/one electron/outer shell becomes complete [1]  
ALLOW: (for 1 mark) potassium loses two electrons + chlorine gains two electrons  
ALLOW: e.g. 2.8.8.1 → 2.8.8 for first mark  
Any indication of sharing electrons = 0

**Total = 12**



Page 4	Mark Scheme	Syllabus	Paper
	Chemistry - June 2004	0620	02

- 6 (a) carbon monoxide [1]
- (b) iron oxide loses oxygen/it loses oxygen/oxidation number of iron decreases [1]  
ALLOW: iron gains electrons  
Answer must refer to the iron/iron oxide - therefore:  
NOT: carbon monoxide gains oxygen  
NOT: oxygen lost in the reaction  
NOT: iron loses oxygen
- (c) 3; 2 (one mark each) [2]
- (d) (i) oxidise the impurities/oxidise Si or P or C/burn off the impurities [1]  
NOT: get rid of impurities  
NOT: slag formation
- (ii) exothermic [1]
- (iii) is/floats above the molten iron [1]
- (iv) calcium oxide [1]
- (v) stronger/harder/not brittle/less easily corroded ORA e.g. iron rusts [1]  
NOT: less corrosive
- (e) any 3 of:  
high melting/boiling points;  
have coloured compounds (NOT: they are coloured);  
have high densities;  
form complex ions;  
elements/compounds are good catalysts;  
form ions with different charges/variable oxidation states [3]
- (f) alloys [1]

**Total = 13**

**Grand Total = 80**

June 2004

**INTERNATIONAL GCSE**

**MARK SCHEME**

**MAXIMUM MARK: 80**

**SYLLABUS/COMPONENT: 0620/03**

**CHEMISTRY  
Extended**



<b>Page 1</b>	<b>Mark Scheme</b>	<b>Syllabus</b>	<b>Paper</b>
	<b>Chemistry – June 2004</b>	<b>0620</b>	<b>3</b>

- When the name of a chemical is demanded by the question, a **correct** formula is usually acceptable. When the formula is asked for, the name is not acceptable.
- When a word equation is required a **correct** symbol equation is usually acceptable. If an equation is requested then a word equation is not usually acceptable.
- An incorrectly written symbol, e.g. NA **or** CL, should be penalised once in a question.

In the mark scheme if a word **or** phrase is underlined it (**or** an equivalent) is required for the award of the mark.

(.....) is used to denote material that is not specifically required.

**OR** designates alternative and independent ways of gaining the marks for the question.

**or** indicates different ways of gaining the same mark.

**COND** indicates that the award of this mark is conditional upon a previous mark being gained.

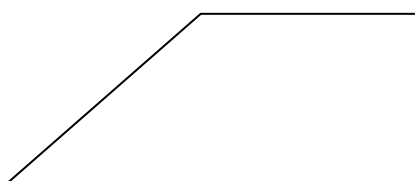
- Unusual responses which include correct Chemistry that answers the question should always be rewarded - even if they are not mentioned in the mark scheme.
- All the candidate's work must show evidence of being marked by the examiner.

Page 2	Mark Scheme	Syllabus	Paper
	Chemistry – June 2004	0620	3

- |    |     |       |  |                     |
|----|-----|-------|--|---------------------|
| 1. | (a) | (i)   | portable   | [1]                 |
|    |     | (ii)  | oxygen <b>or</b> air   | [1]                 |
|    | (b) | (i)   | both have four outer <b>or</b> valency electrons<br>need to share four more<br><b>or</b> need four more to complete energy level<br><b>NOT</b> four bonds  | [1]<br>[1]          |
|    |     | (ii)  | hard<br>brittle<br>high melting <b>or</b> boiling point<br>poor conductor of electricity <b>or</b> semi-conductor<br>any <b>TWO</b><br><b>NOT</b> insoluble in water, <b>NOT</b> tough<br><b>NOT</b> appearance  | [2]                 |
|    |     | (iii) | germanium <b>or</b> carbon<br><b>NOT</b> graphite  | [1]                 |
|    | (c) | (i)   | correctly balanced   | [1]                 |
|    |     | (ii)  | lost oxygen<br><b>or</b> decrease in oxidation number<br><b>NOT</b> accepts electrons unless valid explanation   | [1]                 |
|    |     | (iii) | 4 oxygen atoms around 1 silicon atom<br>2 silicon atoms around 1 oxygen<br>tetrahedral <b>or</b> diagram that looks tetrahedral<br>If some wrong chemistry, such as ionic MAX<br>2/3   | [1]<br>[1]<br>[1]   |
|    |     |       |  | <b>TOTAL = [12]</b> |
| 2. | (a) | (i)   | USA <b>or</b> Texas <b>or</b> Poland <b>or</b> Mexico <b>or</b> Japan <b>or</b> Ethiopia<br>Australia <b>or</b> Sicily<br>accept other sources of sulphur eg petroleum<br><b>or</b> natural gas <b>or</b> metal sulphides <b>or</b> volcanoes<br><b>NOT</b> coal, <b>NOT</b> underground | [1]                 |
|    |     | (ii)  | Preserving food <b>or</b> bleaching <b>or</b> sterilising <b>or</b><br>disinfecting <b>or</b> making paper <b>or</b> bleaching wood pulp<br><b>or</b> wine <b>or</b> jam <b>or</b> fumigation <b>or</b> making paper<br><b>NOT</b> making wood pulp                                      | [1]                 |
|    |     | (iii) | <u>burnt/roast in oxygen <b>or</b> air</u>   | [1]                 |
|    |     | (iv)  | vanadium(V) oxide <b>or</b> vanadium oxide <b>or</b> platinum<br>ignore oxidation state of vanadium  | [1]                 |
|    |     | (v)   | Increase temperature (increases rate) but reduces yield<br>catalyst only increases rate <b>or</b> a catalyst does not<br>influence position of equilibrium<br><b>NOT</b> a definition of a catalyst  | [1]<br>[1]          |
|    |     | (vi)  | sulphur trioxide + sulphuric acid = oleum<br>correct symbol equation acceptable  | [1]                 |
|    |     | (vii) | $H_2S_2O_7 + H_2O = 2H_2SO_4$  | [1]                 |

Page 3	Mark Scheme	Syllabus	Paper
	Chemistry – June 2004	0620	3

- (b) (i) potassium [1]
- (ii) ammonium sulphate [1]
- (iii)  $\text{Ca}_3(\text{PO}_4)_2$  [1]
- $\text{Ca}(\text{H}_2\text{PO}_4)_2$  [1]
- (iv) only acceptable responses are:  
accepts a proton [2]  
accepts  $\text{H}^+$  [1] only
- TOTAL = [14]**
3. (a) dissolved **or** solution in water [1]  
**NOT** aqueous **NOT** soluble in water  
l liquid and g gas [1]
- (b) 6 electrons in bond between two nitrogen atoms [1]  
2 electrons on each nitrogen [1]  
ignore any coding of electrons with dots **or** crosses
- (c) (i) decreases **or** reaction stops **or** rate becomes zero [1]
- (ii) concentration **or** number of effective collisions  
decreases [1]  
used up **or** less chemical **or** less collisions etc [1] only
- (iii) greater initial slope [1]  
same final point [1]  
as long as new curve touches the original curve near  
the top allocate the mark
- (iv) greater surface area [1]
- TOTAL = [10]**
- 4 (a) (i) Named soluble zinc salt [1]  
corresponding sodium salt [1]  
If hydroxide **or** oxide then 0/2
- (ii) Correct equation [2]  
not balanced [1] only
- (iii) Correct equation [2]
- (b) (i)  $\text{Fe}^{3+} + 3\text{OH}^- = \text{Fe}(\text{OH})_3$  [1]
- (ii) Max at  $8\text{cm}^3$  [1]  
Same shape of graph



Just the above shape, the height of the precipitate and the volume of sodium hydroxide are irrelevant [1]

Page 4	Mark Scheme	Syllabus	Paper
	Chemistry – June 2004	0620	3

(iii) Maximum then height of precipitate decreases [1]  
**or** graph slopes down to x axis **or** comes to zero

hydroxide dissolves in excess **or** it is amphoteric [1]

**TOTAL = [11]**

5. (a) Has to be three different uses.

any use that depends on malleability **or** ductility-  
jewellery, pipes, wires, sheets, roofing, ornaments [1]  
**NOT** that it is malleable **or** ductile

electrical wires **or** cooking utensils **or** electrodes [1]  
(good) conductor

making alloys **or** named alloy [1]

(b) (i)  $\text{Cu}^{2+} + 2\text{e} = \text{Cu}$  [1]

(ii) gas is oxygen [1]

(copper(II) sulphate) changes to sulphuric acid  
**or** copper ions removed from solution [1]

(c) (i) copper atoms - electrons = copper ions [1]  
accept correct symbol equation

(ii) concentration of copper ions does not change **or** [1]  
amount **or** number of copper ions does not change

copper ions are removed and then replaced [1]  
**or** copper is transferred from anode to cathode

(iii) refining copper **or** plating (core) [1]  
**or** extraction of boulder copper

**TOTAL = [10]**

6. (a) (i) correct repeat unit [1]

**COND** evidence of polymer chain [1]

(ii) glucose **or** maltose [1]

(iii) addition (polymerisation) **or** no other product [1]  
except polymer

condensation (polymerisation) **or** polymer [1]  
and water

(b) (i) sodium hydroxide [1]

**COND** ammonia **or** alkaline gas **or** litmus red to blue [1]  
If aluminium added  $w_c = 0$

Page 5	Mark Scheme	Syllabus	Paper
	Chemistry – June 2004	0620	3

- (ii) measure pH [1]  
more than 1 and less than 7 **or**  
correct colour eg orange **or** yellow **NOT** red  
**NOT** green [1]  
**OR** add magnesium **or** calcium carbonate [1]  
weak acid reacts slowly
- (c) (i) ethyl acrylate [1]  
ester **or** alkene [1]
- (ii) brown to colourless (**NOT** clear) [1]  
correct formula for acid **NOT** ester [1]
- TOTAL = [13]**
- 7 (a) Avogadro's Number of particles  
**or** formula mass in grams  
**or**  $6 \times 10^{23}$  particles accept atoms, ions and molecules  
**or** as many particles as there are carbon atoms in 12.00g of  $^{12}\text{Ca}$   
ANY one [1]
- (b) (i) moles of Mg =  $3/24 = 0.125$   
moles of  $\text{CH}_3\text{COOH} = 12/60 = 0.200$   
magnesium is in excess  
  
**OR** 3.0g of magnesium react with 15g of acid  
only 12.0 g of acid present  
magnesium is in excess [3]
- (ii) **Mark conseq to (i) but NOT to any simple integer**  
moles of  $\text{H}_2 = 0.1$  [1]
- (iii) **Mark conseq to (ii) but NOT to any simple integer**  
Volume of hydrogen =  $0.1 \times 24$   
=  $2.4 \text{ dm}^3$  [2]
- (c) (i) moles of NaOH =  $25/1000 \times 0.4 = 0.01$  [1]
- (ii) **Mark conseq to (i) but NOT to any simple integer**  
moles of acid =  $0.01/2 = 0.005$  [1]
- (iii) **Mark conseq to (ii) max 10M**  
concentration of acid =  $0.005 \times 1000/20$  [1]  
=  $0.25 \text{ mol/dm}^3$  [1]
- TOTAL = [10]**

**TOTAL for PAPER = [11] + [14] + [10] + [11] + [10] + [13] + [11] = [80]**

**June 2004**

**INTERNATIONAL GCSE**

**MARK SCHEME**

**MAXIMUM MARK: 40**

**SYLLABUS/COMPONENT: 0620/05**

**CHEMISTRY  
Practical**



<b>Page 1</b>	<b>Mark Scheme</b>	<b>Syllabus</b>	<b>Paper</b>
	<b>Chemistry – June 2004</b>	<b>0620</b>	<b>5</b>

<b>1</b>	<b>Table of results</b>			
	Experiment 1			
	Temperature boxes completed		1	
	Increasing		1	
	Comparable to supervisor		1	[3]
	Experiment 2			
	Temperature boxes completed		1	
	Decreasing		1	
	Comparable to supervisor		1	[3]
<b>(a)</b>	All points plotted correctly (-1 for each incorrect)		4	
	Smooth line graphs		2	
	Labelled		1	[7]
<b>(b)</b>	<b>(i)</b>	1. Value from graph 2. Value from graph $\pm 0.25$ } No unit only (1)	1 1	[2]
	<b>(ii)</b>	1. Exothermic 2. Endothermic	1 1	[2]
<b>(c)</b>	Fizz/bubbles/effervescence		1	
	Solid disappears		1	[2]
<b>(d)</b>	Carbonate		1	
	Fizz with acid or similar		1	[2]
<b>(e)</b>	Solid <b>A</b> – value from table/room temperature $\pm 3^{\circ}\text{C}$		1	
	Solid <b>B</b> – value from table/room temperature		1	
	Reaction finished		1	[3]
			<b>Sub Total</b>	<b>[24]</b>
<b>2</b>	<b>(a)</b>	White	1	[1]
	<b>(c)</b>	<b>(i)</b>		
		White	1	
		Precipitate	1	[2]
		Excess – no change	1	[1]
	<b>(ii)</b>	No precipitate/change	1	[1]
	<b>(iii)</b>	Paper goes blue	1	
		Fizz/bubbles etc	1	
		Reference to smell	1	[3]
	<b>(iv)</b>	pH greater than 7	1	[1]
	<b>(v)</b>	Milky/cloudy	1	[1]
<b>(d)</b>	Calcium		1	[1]
<b>(e)</b>	Ammonia		1	[1]

<b>Page 2</b>	<b>Mark Scheme</b>	<b>Syllabus</b>	<b>Paper</b>
	<b>Chemistry – June 2004</b>	<b>0620</b>	<b>5</b>

<b>(f)</b>	Limewater	1	<b>[2]</b>
	Carbon dioxide	1	
<b>(g)</b>	Nitrate	1	<b>[2]</b>
	Hydroxide	1	
		<b>Sub Total</b>	<b>[16]</b>
		<b>Total</b>	<b>[40]</b>

**June 2004**

**INTERNATIONAL GCSE**

**MARK SCHEME**

**MAXIMUM MARK: 60**

**SYLLABUS/COMPONENT: 0620/06**

**CHEMISTRY**  
**Alternative to Practical**





Page 2	Mark Scheme	Syllabus	Paper
	IGCSE – June 2004	0620	6

	(d) (i)	22°C	} No units only (1)	1	
		21°C		1	
	(ii)	Reference to room temperature/reaction finished		1	[3]
5	(a) (i)	White Precipitate		1 1	[2]
		No change/white precipitate/insoluble in excess		1	[1]
	(ii)	No/thin precipitate/no reaction		1	[1]
	(b)	Ammonia		1	[1]
	(c)	Reference to limewater/test for carbon dioxide		1	[1]
	(d)	Nitrate Alkali/hydroxide/oxide		1 1	[2]
6	(a)	Indication of copper oxide		1	[1]
	(b)	Black to red/pink/brown		1 1	[2]
	(c)	To cool/condense Steam/water		1 1	[2]
7	(a)	Anhydrous copper sulphate/cobalt chloride Goes blue/pink in water, no change for ethanol		1 1	[2]
	(b)	Add indicator/named indicator or CO <sub>3</sub> <sup>2-</sup> /Mg Turns red/correct colour in acid, no change for sodium sulphate		1 1	[2]
	(c)	Add silver nitrate White precipitate with hydrochloric acid, no change with nitric acid		1 1	[2]
8		Add known mass of manganese oxide		1	
		To (measured volume of) hydrogen peroxide		1	
		Bubbles		1	
		Test gas with glowing splint		1	
		Result		1	
		Filter		1	
		Dry solid		1	
		Reweigh and compare		1	
		<b>(max 6)</b>			[6]
				<b>Total for Paper</b>	<b>[60]</b>